

Chromatography and analysis (CR6) – Revision Checklist

I can...	Lesson	Revised
Define a pure substance	CR6 LE3	
Define a mixture.	CR6 LE3	
Describe the difference between a pure substance and a mixture.	CR6 LE3	
Use particle model diagrams for pure substances and mixtures	CR6 LE3	
Explain why substances in a mixture retain their chemical properties.	CR6 LE3	
Describe how melting point data can be used to identify a pure substance.	CR6 LE3	
Describe how boiling point data can be used to identify a pure substance.	CR6 LE3	
Use melting point data to distinguish between pure and impure substances.	CR6 LE3	
Use boiling point data to distinguish between pure and impure substances.	CR6 LE3	
Explain the difference between the everyday meaning of pure and the chemistry definition of pure.	CR6 LE3	
Define a formulation.	CR6 LE3	
Give examples of common household formulations	CR6 LE3	
Explain why components in formulations are measured carefully.	CR6 LE3	
Identify a formulation from given information.	CR6 LE3	
Recognise examples of products that are formulations.	CR6 LE3	
State the purpose of chromatography.	CR6 LE4	
Define the stationary phase in chromatography.	CR6 LE4	
Define the mobile phase in chromatography.	CR6 LE4	
Explain how substances are separated in chromatography.	CR6 LE4	
Explain how paper chromatography works.	CR6 LE4	
Define R _f value.	CR6 LE4	
Calculate R _f values from chromatograms.	CR6 LE4	
Interpret chromatograms.	CR6 LE4	
Use R _f values to help identify substances.	CR6 LE4	
Describe how to carry out paper chromatography	CR6 LE5	
Identify sources of error in paper chromatography	CR6 LE5	
Suggest how chromatography can distinguish pure substances from mixtures.	CR6 LE5	
Describe the test and positive result hydrogen gas.	CR6 LE1	
Describe the test and positive result oxygen gas.	CR6 LE2	
Describe the test and positive result carbon dioxide gas.	CR6 LE1	
Describe the test and positive result chlorine gas.	CR6 LE2	